

WORKSHEET -EQUATIONS

NAME _____
Last First

A) Write balanced chemical equations for the reactions given below .

B) Classify the reaction type as : a) combination, b) decomposition, c) single-replacement, d) double-displacement (Metathesis, or ion exchange), or e) combustion.

_____ 1. Calcium chloride and silver nitrate

_____ 2. Sulfur dioxide and water forming sulfurous acid.

_____ 3. Zinc and cupric nitrate.

_____ 4. Oxalic acid, $\text{H}_2\text{C}_2\text{O}_4$, and sodium hydroxide.

_____ 5. Decomposition of cupric hydroxide.

_____ 6. Ferric chloride and potassium phosphate.

_____ 7. Burning of butane, C_4H_{10} .

_____ 8. Decomposition of sodium chloride.

_____ 9. Magnesium and hydrochloric acid.

_____ 10. Chlorine and barium bromide.

_____ 11. Decomposition of magnesium carbonate.

_____ 12. Decomposition of hydrogen peroxide.

(OVER)

_____ 13. Potassium oxide and carbon dioxide

_____ 14. Sodium and water.

_____ 15. Sodium oxide and water.

_____ 16. Decomposition of sodium nitrate.

_____ 17. Calcium oxide and water.

_____ 18. Potassium iodide and lead nitrate.

_____ 19. Sulfuric acid and calcium hydroxide.

_____ 20. Diphosphorus pentoxide and water forming phosphoric acid.

_____ 21. Decomposition of calcium sulfite.

_____ 22. Magnesium oxide and sulfur dioxide.

_____ 23. Calcium and sulfuric acid.

_____ 24. Magnesium oxide and water.

_____ 25) Decomposition of sodium peroxide.

_____ 26) Decomposition of potassium chlorate.

