

# Net-Ionic Equations

## CHEMISTRY 110

Name \_\_\_\_\_  
Last First

- Write a balanced equation for the reactants given.
- Include the physical states for all reagents: Assume that all reactions are in water.
- Write a total ionic equation
- Write the net-ionic equation

*hint: Use solubility rules, activity tables, and tables for strong bases and acids to write the equations!*

**1]** Nickel (III)chloride + potassium phosphate -->

Molecular equation \_\_\_\_\_

Total ionic \_\_\_\_\_

Net ionic \_\_\_\_\_

**2]** Aluminum + Hydrobromic acid -->

Molecular equation \_\_\_\_\_

Total ionic \_\_\_\_\_

Net ionic \_\_\_\_\_

**3]** Sodium carbonate + silver nitrate -->

Molecular equation \_\_\_\_\_

Total ionic \_\_\_\_\_

Net ionic \_\_\_\_\_

**4]** Silver hydroxide + carbonic acid -->

Molecular equation \_\_\_\_\_

Total ionic \_\_\_\_\_

Net ionic \_\_\_\_\_

**5]** Zinc + Stannic nitrate -->

Molecular equation \_\_\_\_\_

Total ionic \_\_\_\_\_

Net ionic \_\_\_\_\_

**6]** Potassium carbonate + aluminum nitrate -->

Molecular equation \_\_\_\_\_

Total ionic \_\_\_\_\_

Net ionic \_\_\_\_\_

**B**

**7]** Cesium cyanide + hydrochloric acid -->

Molecular equation \_\_\_\_\_

Total ionic \_\_\_\_\_

Net ionic \_\_\_\_\_

**8]** Barium hydroxide + cupric acetate -->

Molecular equation \_\_\_\_\_

Total ionic \_\_\_\_\_

Net ionic \_\_\_\_\_

**9]** Chromium (III) chloride + sodium nitrate -->

Molecular equation \_\_\_\_\_

Total ionic \_\_\_\_\_

Net ionic \_\_\_\_\_

**10]** Iron + nickel(III) iodide -->

Molecular equation \_\_\_\_\_

Total ionic \_\_\_\_\_

Net ionic \_\_\_\_\_

**11]** Sodium hypiodite + ammonium dichromate -->

Molecular equation \_\_\_\_\_

Total ionic \_\_\_\_\_

Net ionic \_\_\_\_\_

**12]** Acetic acid + Chromic hydroxide -->

Molecular equation \_\_\_\_\_

Total ionic \_\_\_\_\_

Net ionic \_\_\_\_\_

**13]** Ferrous nitrate + sodium sulfide -->

Molecular equation \_\_\_\_\_

Total ionic \_\_\_\_\_

Net ionic \_\_\_\_\_

**B**

**14]** Phosphoric acid + cobalt (II) bromide -->

Molecular equation \_\_\_\_\_

Total ionic \_\_\_\_\_

Net ionic \_\_\_\_\_

**15]** Sodium sulfide + lead (II) nitrate-->

Molecular equation \_\_\_\_\_

Total ionic \_\_\_\_\_

Net ionic \_\_\_\_\_