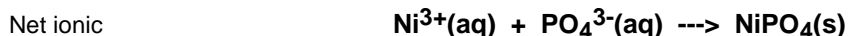
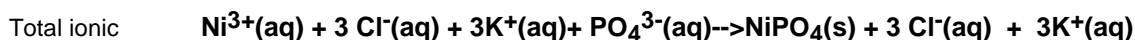
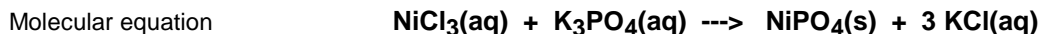


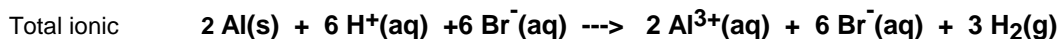
Net-Ionic Equations Chemistry 110

- a) Write a balanced equation for the reactants given.
b) Include the physical states for all reagents: Assume that all reactions are in water.
c) Write a total ionic equation
d) Write the net-ionic equation
hint: Use solubility rules, activity tables, and tables for strong bases and acids to write the equations!

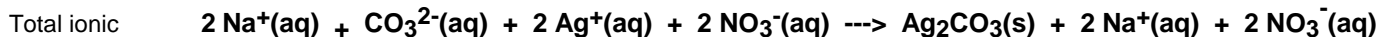
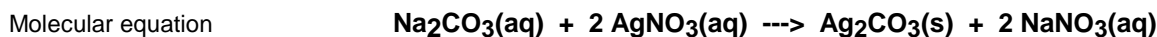
1] Nickel (III)chloride + potassium phosphate -->



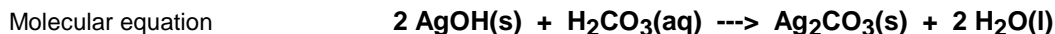
2] Aluminum + Hydrobromic acid -->



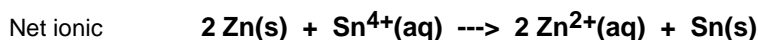
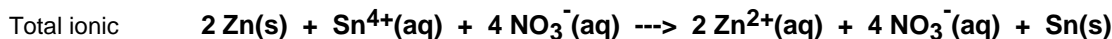
3] Sodium carbonate + silver nitrate -->



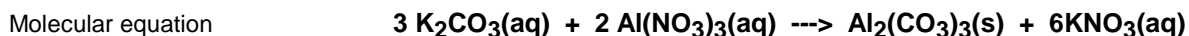
4] Silver hydroxide + carbonic acid -->

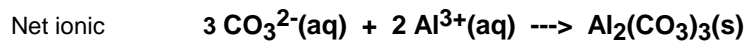


5] Zinc + Stannic nitrate -->

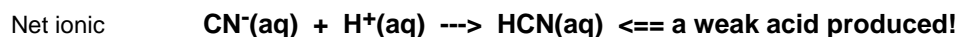
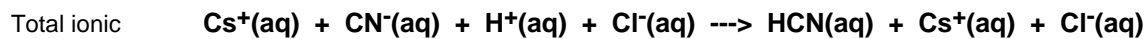


6] Potassium carbonate + aluminum nitrate -->

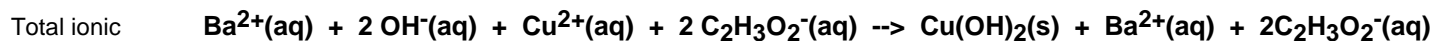
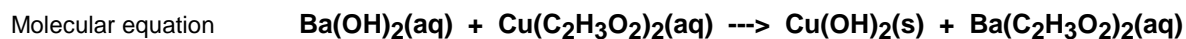




7] Cesium cyanide + hydrochloric acid -->

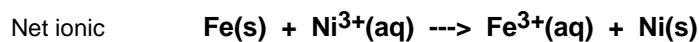
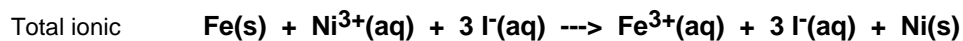
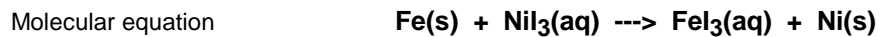


8] Barium hydroxide + cupric acetate -->

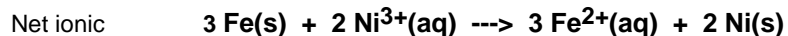
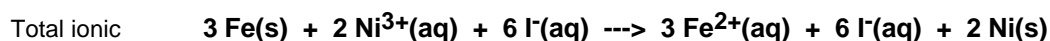
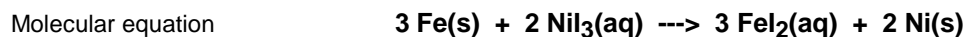


9] Chromium (III) chloride + sodium nitrate --> **NR**

10] Iron + nickel(III) iodide -->

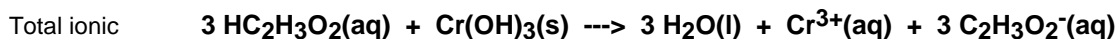


OR



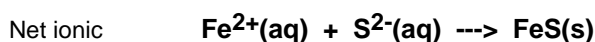
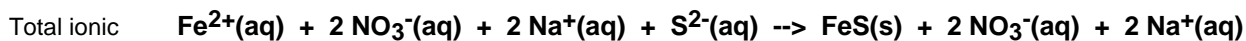
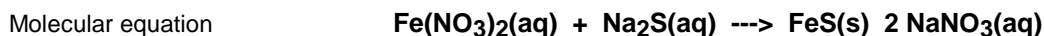
11] Sodium hypoiodite + ammonium dichromate --> **NR**

12] Acetic acid + Chromic hydroxide -->

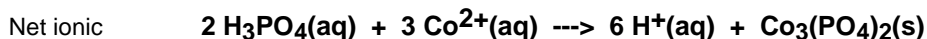
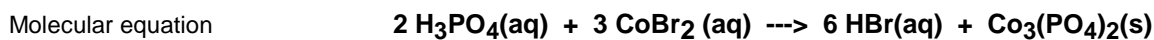


Net ionic **SAME AS TOTAL**

13] Ferrous nitrate + sodium sulfide -->



14]. Phosphoric acid + cobalt (II) bromide -->



15] Sodium sulfide + lead (II) nitrate-->

