

Nomenclature Chemistry 110

1. Write the chemical formulas for the compounds formed between the following ions:

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|---------------------------------------|---|
| a. Potassium-----bromide | <u> KBr </u> |
| b. Ferric ion----Sulfide | <u> Fe₂S₃ </u> |
| c. Tin(IV)----Phosphide | <u> Sn₃P₄ </u> |
| d. Manganese (III)-----Sulfite | <u> Mn₂(SO₃)₃ </u> |
| e. Chromic---Oxalate | <u> Cr₂(C₂O₄)₃ </u> |
| f. Zinc-----Bromite | <u> Zn(BrO₂)₂ </u> |
| g. Aluminum----iodate | <u> Al(IO₃)₃ </u> |
| h. Cobaltous-----Bicarbonate | <u> Co(HCO₃)₂ </u> |
| i. Mercury (I)-----Nitrate | <u> Hg₂(NO₃)₂ </u> |
| j. Barium-----mono hydrogen phosphate | <u> BaHPO₄ </u> |

2. Predict the formula for the following compounds:

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|---------------------------|---|---------------------------|--|
| a. Strontium chloride | <u> SrCl₂ </u> | b. Sulfurous Acid | <u> H₂SO₃ </u> |
| c. Diarsenic Pentasulfide | <u> As₂S₅ </u> | d. Cupric Hypochlorite | <u> Cu(ClO)₂ </u> |
| e. Nickel (II) Phosphate | <u> Ni₃(PO₄)₂ </u> | f. Zinc Nitride | <u> Zn₃N₂ </u> |
| g. Silver Dichromate | <u> Ag₂Cr₂O₇ </u> | h. Potassium Permanganate | <u> KMnO₄ </u> |
| i. Carbon tetrachloride | <u> CCl₄ </u> | j. Periodic Acid | <u> HIO₄ (aq) </u> |

3. Name the following compounds:

- | | |
|--|---|
| a. P ₄ O ₆ | <u> Tetraphosphorus hexaoxide [hexoxide] </u> |
| b. HBrO ₄ (aq) | <u> Perbromic acid </u> |
| c. HBr(aq) | <u> Hydrobromic acid </u> |
| d. HgCrO ₄ | <u> Mercury(II) chromate [Mercuric] </u> |
| e. Sn(SO ₃) ₂ | <u> Tin(IV) sulfite [Stannic] </u> |
| f. OF ₂ | <u> Oxygen difluoride </u> |
| g. AlN | <u> Aluminum nitride </u> |
| h. K ₂ O ₂ | <u> Potassium peroxide </u> |
| i. H ₂ CO ₃ (aq) | <u> Carbonic acid </u> |
| j. SiO ₄ | <u> Silicon tetraoxide [tetroxide] </u> |