

WORKSHEET 8**CHEMISTRY 110**

Name _____

last

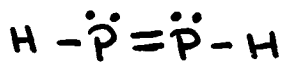
first

Problem sets are due within the first five minutes of lecture on the due date.
No late work will be accepted.

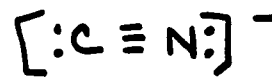
Draw the electron dot structure for the following compounds. You must show all electrons, bonding and nonbonding:

<p>Br₂</p> <p>$\begin{array}{c} \cdot\cdot \\ \cdot\cdot \\ :\text{Br} - \text{Br}: \\ \cdot\cdot \\ \cdot\cdot \end{array}$</p>	<p>PCl₃</p> <p>$\begin{array}{c} \cdot\cdot & & \cdot\cdot \\ \cdot\cdot & & \cdot\cdot \\ :\text{Cl} - \text{P} - \text{Cl}: \\ \\ \cdot\cdot & & \cdot\cdot \\ \cdot\cdot & & \cdot\cdot \\ :\text{Cl}: \\ \cdot\cdot & & \cdot\cdot \end{array}$</p>
<p>H₂Se</p> <p>$\begin{array}{c} \text{H} \\ \\ \text{H} - \text{Se}: \\ \cdot\cdot \end{array}$</p>	<p>SCl₂</p> <p>$\begin{array}{c} \cdot\cdot \\ \cdot\cdot \\ :\text{Cl}: \\ \\ \cdot\cdot & & \cdot\cdot \\ \cdot\cdot & & \cdot\cdot \\ :\text{S} - \text{Cl}: \\ \cdot\cdot & & \cdot\cdot \end{array}$</p>
<p>NO₂F</p> <p>$\begin{array}{c} \cdot\cdot \\ \cdot\cdot \\ :\text{F}: \\ \\ \cdot\cdot & & \cdot\cdot \\ \cdot\cdot & & \cdot\cdot \\ :\text{O} - \text{N} = \text{O}: \\ \cdot\cdot & & \cdot\cdot \end{array}$</p>	<p>PO_3^{3-}</p> <p>$\left[\begin{array}{c} \cdot\cdot \\ \cdot\cdot \\ :\text{O}: \\ \\ \cdot\cdot & & \cdot\cdot \\ \cdot\cdot & & \cdot\cdot \\ :\text{O} - \text{P} - \text{O}: \\ \cdot\cdot & & \cdot\cdot \end{array} \right]^{3-}$</p>

HPPH
(P₂H₂)



CN⁻

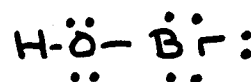


charge & brackets must be shown.

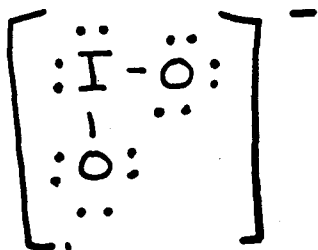
SO₂



BrO⁻

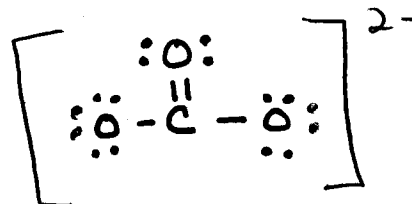


IO₂⁻

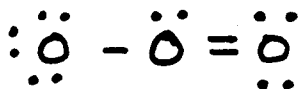


charge & brackets must be shown.

CO₃²⁻



OO
(O₃)



COCl₂

