

		<b>Is it a Buffer?</b>		
a.	Weak Acid H F	+	its salt Na F	
b.	Weak Base NH <sub>4</sub> OH	+	its salt H <sub>4</sub> Cl	
	<u>HC<sub>2</sub>H<sub>3</sub>O<sub>2</sub></u> WA	+	LiC <sub>2</sub> H <sub>3</sub> O <sub>2</sub>	YES
	<u>HC<sub>7</sub>H<sub>12</sub>O<sub>6</sub></u> WA	+	K C <sub>7</sub> H <sub>12</sub> O <sub>6</sub>	YES
	HNO <sub>2</sub> WA	+	HN0 <sub>3</sub>	NO
	<u>HClO<sub>3</sub></u> WA	+	Ca (ClO <sub>3</sub> ) <sub>2</sub>	YES
	NH <sub>4</sub> OH WB	+	(NH <sub>4</sub> ) <sub>2</sub> SO <sub>4</sub>	YES
	NH <sub>4</sub> Cl NO Salt	+	LiC <sub>2</sub> H <sub>3</sub> O <sub>2</sub> salt	

What is the pH of a 0.001 M KCl solution?

pH= 7

What has a higher pH

0.001 M HCl

or 0.001 M KOH

H<sup>+</sup> = 10<sup>-3</sup> strong  
Acid

OH<sup>-</sup> = 10<sup>-3</sup> strong  
base

pH = 3

H<sup>+</sup> = 10<sup>-11</sup>

ph= 11