

## EXERCISE 9

Chem 100

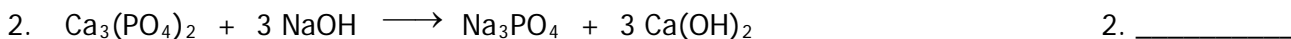
(Due date \_\_\_\_\_)

10 points

Name \_\_\_\_\_  
( last) (first)

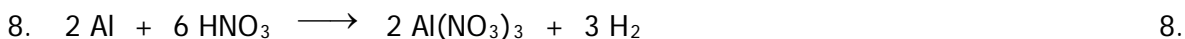
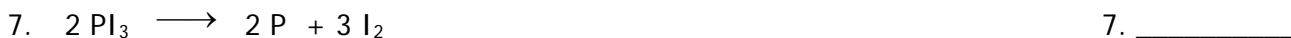
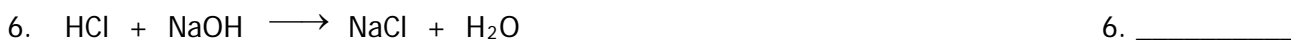
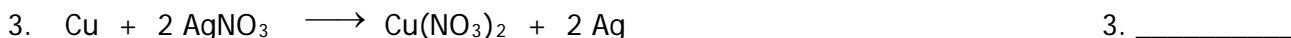
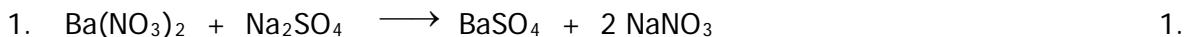
Lecture Section # \_\_\_\_\_ Instructor \_\_\_\_\_

A. For each of the following reactions, write a "**B**" if the equation is balanced and a "**U**" if the equation is unbalanced.



B. Reaction Type: Write the correct *letter* in the blank at the right.

(A) Combination (B) Single Replacement (C) Decomposition (D) Double Replacement





(OVER)

C. For each of the following DOUBLE REPLACEMENT reactions, *circle* the **PRODUCT** which causes the reaction to occur (go to completion). Then write the correct letter that corresponds to the product you circled in the blank at the right - that is - choose from the following:

(A) weak acid

(D) insoluble gas

(B) weak base

(E) water

(C) insoluble salt (precipitate)

(F) no reaction

