

EXERCISE 2

Chem 10

Due in lab _____

10 points

Name Key
(last) (first)

Lab Section # _____ Lab Instructor _____

- A. Elements. Complete the following table. In the column labeled "family", write the common name of the family (alkali metal, alkaline earth metal, transition metal, halogen, or noble gas) to which the element belongs. If the element's family does not have a common name, write only the group number (for example, IIIA). Get atomic masses from the periodic chart that was handed out to you in class.

SYMBOL	NAME OF ELEMENT	ATOMIC NUMBER	ATOMIC MASS	METAL OR NONMETAL	FAMILY (Group)
F	fluorine	9	19.0	non	halogen
Cu	copper	29	63.5	met	transition metal
Si	silicon	14	28.1	non	IVA
Ca	calcium	20	40.1	met	alkaline earth metal
Fe	iron	26	55.8	met	transition metal
Ar	argon	18	39.9	non	noble gas
Al	aluminum	13	27.0	met	III A
Na	sodium	11	23.0	met	alkali metal
S	sulfur	16	32.1	non	VI A
Au	gold	79	197.0	met	transition metal

- B. Diatomic Elements. Write the names of any four (4) of the seven diatomic elements and the corresponding formulas which show the molecular composition of each.

	<u>NAME</u>	<u>FORMULA</u>
1.	<u>hydrogen</u>	<u>H₂</u>
	<u>nitrogen</u>	<u>N₂</u>
2.	<u>Oxygen</u>	<u>O₂</u>
	<u>fluorine</u>	<u>F₂</u>
3.	<u>chlorine</u>	<u>Cl₂</u>
	<u>bromine</u>	<u>Br₂</u>
4.	<u>Iodine</u>	<u>I₂</u>

- C. Period & Family. Identify each set of elements as members of the same family (group) or the same period of elements by writing "FAMILY" or "PERIOD" in the blanks at the right.

- | | |
|---------------|-------------------------|
| 1. Mg, Al, Cl | 1. <u>family period</u> |
| 2. Be, Ca, Ba | 2. <u>family</u> |
| 3. F, Br, I | 3. <u>family</u> |
| 4. B, Al, Ga | 4. <u>family</u> |