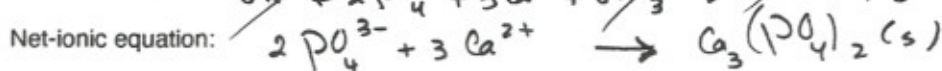
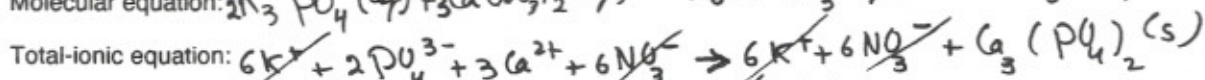
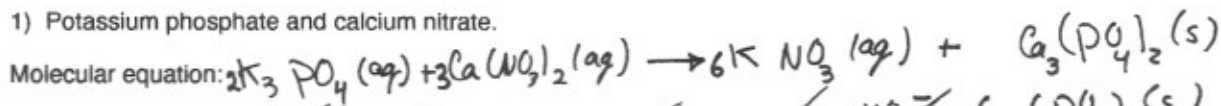


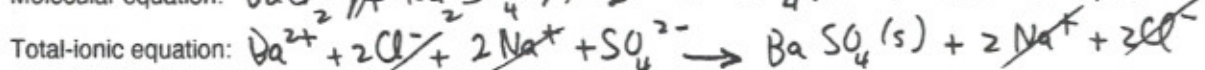
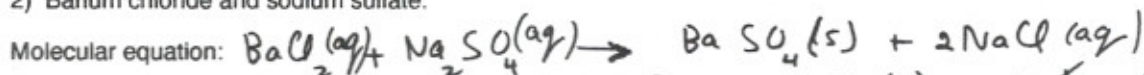
(Any soluble substance given below is in aqueous solution.)

Write molecular, total-ionic and net-ionic equations for the following:

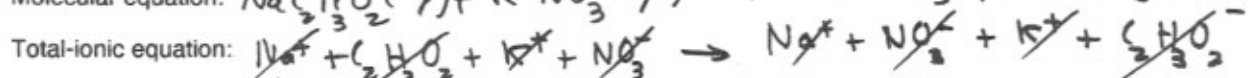
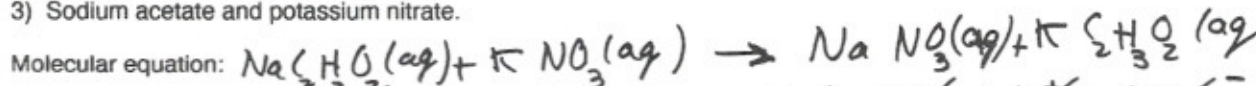
1) Potassium phosphate and calcium nitrate.



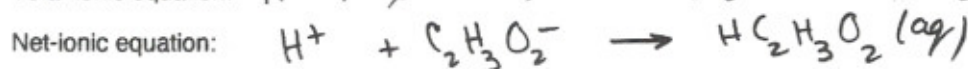
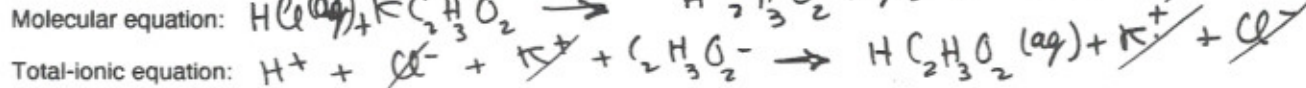
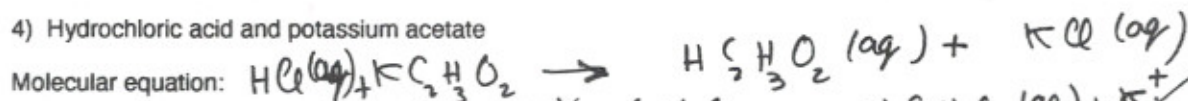
2) Barium chloride and sodium sulfate.



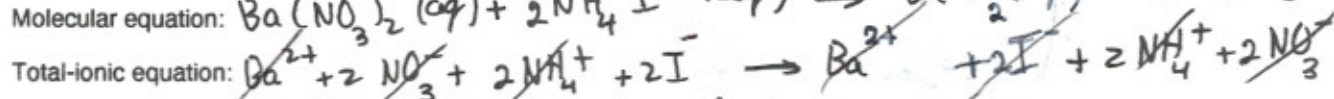
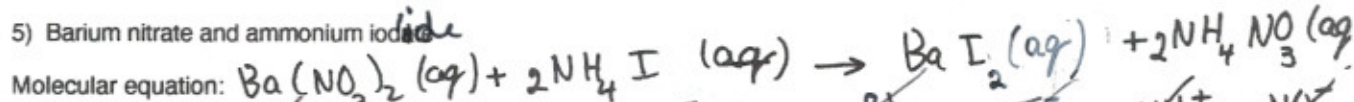
3) Sodium acetate and potassium nitrate.



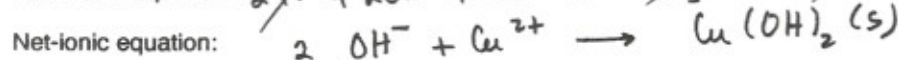
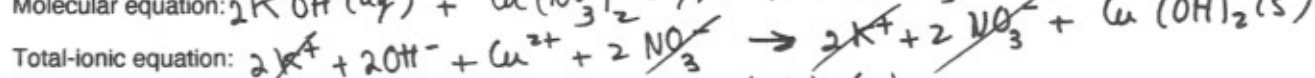
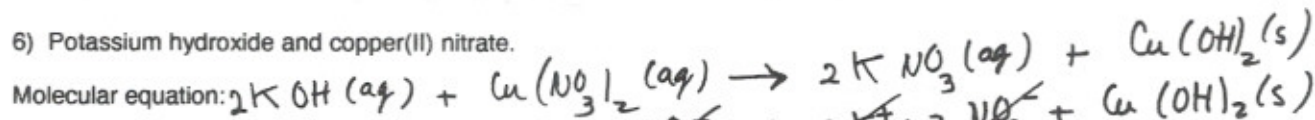
4) Hydrochloric acid and potassium acetate



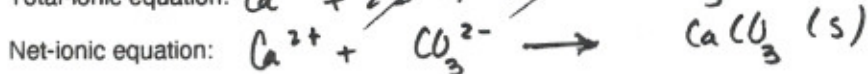
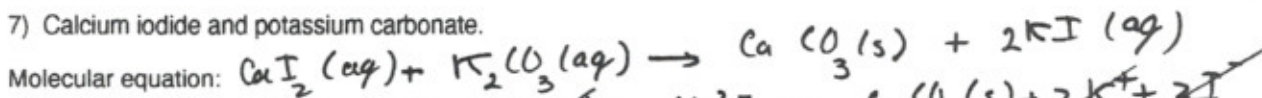
5) Barium nitrate and ammonium iodide



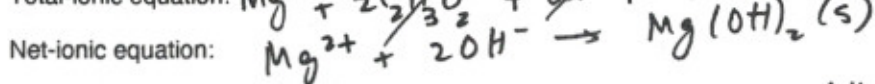
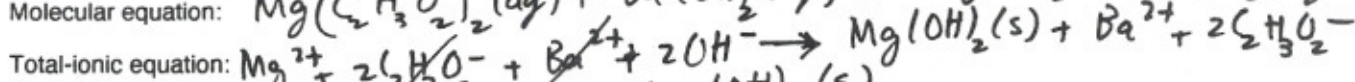
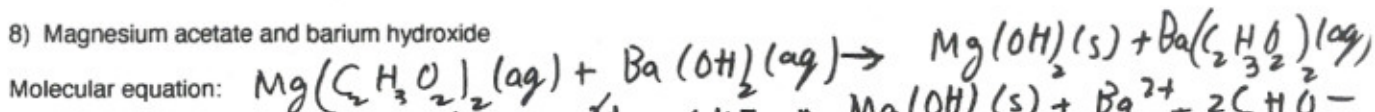
6) Potassium hydroxide and copper(II) nitrate.



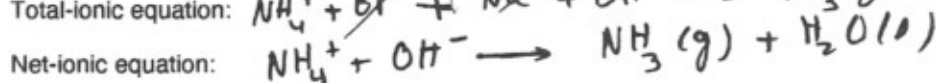
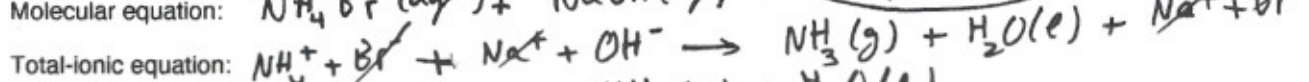
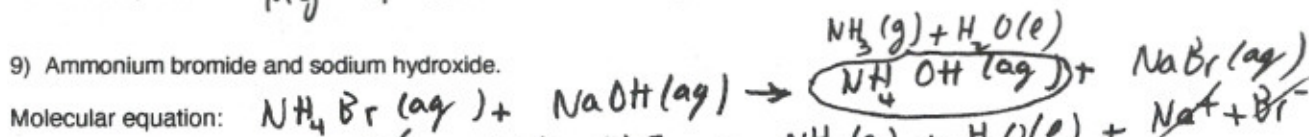
7) Calcium iodide and potassium carbonate.



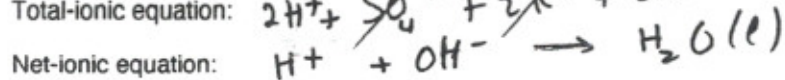
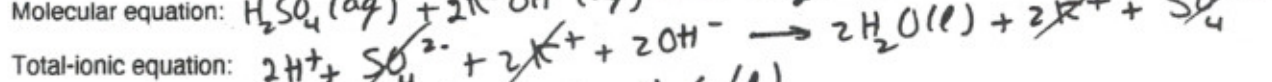
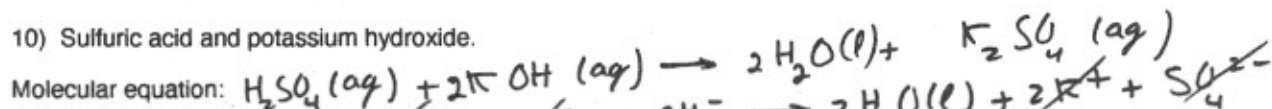
8) Magnesium acetate and barium hydroxide



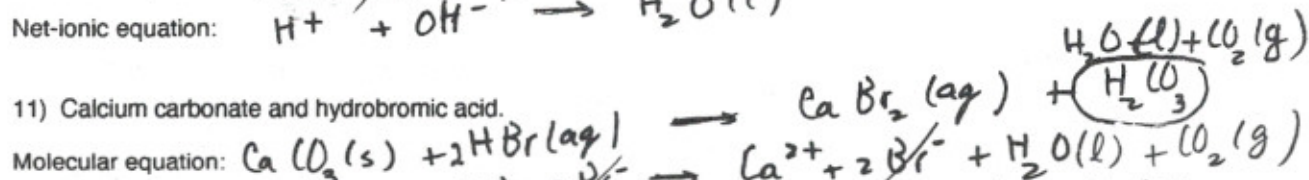
9) Ammonium bromide and sodium hydroxide.



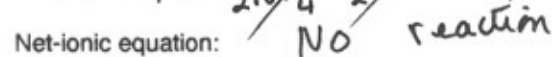
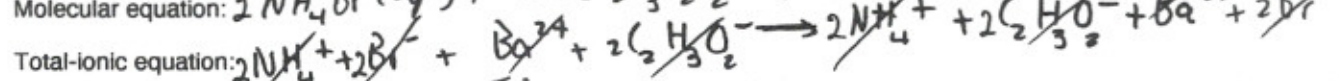
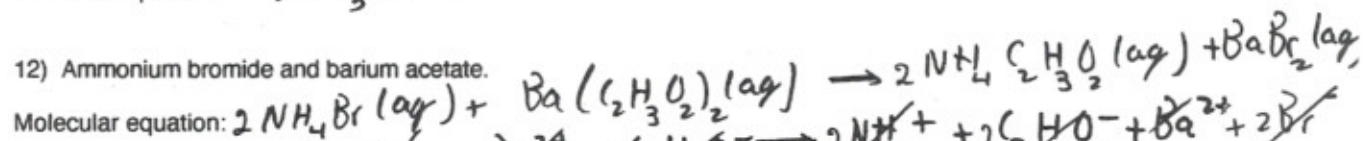
10) Sulfuric acid and potassium hydroxide.



11) Calcium carbonate and hydrobromic acid.



12) Ammonium bromide and barium acetate.



13) Barium hydroxide and ammonium bromide.

