Redox Equations

Acidic:
$$Cl_2 + Mn^{2+} \rightarrow Mn_2O_3 + Cl^-$$

 $\rightarrow 2Mn^{2+} + 3H_2O + Cl_2 \rightarrow Mn_2O_3 + 6H^+ + 2Cl^-$

Redox equal Half-rxn method

Acidic:
$$Cl_2 + Mn^{2+} \rightarrow Mn_2O_3 + Cl^- \rightarrow 2Mn^{2+} + 3H_2O + Cl_2 \rightarrow Mn_2O_3 + 6H^+ + 2Cl^-$$

Basic: $I^- + Cu^{2+} \rightarrow Cu + IO_3^- \rightarrow 3 Cu^{2+} + I^- + 60H^- \rightarrow 3 Cu + IO_3^- + 3H_2O$

$$Fe^{3+} + Br^{-} \rightarrow Fe^{2+} + Br0^{-}$$

$$\longrightarrow 2Fe^{3+} + Br^{-} + 20H^{-} \rightarrow 2Fe^{2+} + Br0^{-} + H_{2}$$