

Perform the following text book problems in the space provided. Show all work. No credit will be given if no work is shown for problems that require work. These problems are **due within the first 5 minutes** of the beginning of lecture on the due date. **No late work will be accepted.**  
It is suggested that students first work out the assigned problems (found in the syllabus) in which the answers are in the back of the text book.

Show Work:	Answers:
10.74b	
10.74d	
10.82	

Worksheet 13

13.106 pg. 493	
10.88	
10.90	
13.152	

25.6 g of silver nitrite (molar mass=153.9) is reacted with 27.3 grams of potassium carbonate (molar mass=138.0) which produces silver carbonate and potassium nitrite.

a. Write and balance the equation:

b. Calculate the number of grams of silver carbonate (molar mass=275.7) produced.

c. How many grams of silver nitrite will remain at the end of the reaction?

d. How many grams of potassium carbonate will remain at the end of the reaction?

