

.These problems are **due within the first 5 minutes** of the beginning of lecture on the due date. **No late work will be accepted.**

a) Write a balanced equation for the reactants given. b) Include the physical states for all reagents: Assume that all reactions are in water. c) Write a total ionic equation d) Write the net-ionic equation

hint: Use solubility rules, activity tables, and tables for strong bases and acids to write the equations

1] Aluminum + tin (II) nitrate -->
Molecular equation
Total ionic
Net ionic
2] Potassium phosphate + bismuth (V) acetate -->
Molecular equation
Total ionic
Net ionic
3] Magnesium + chromium (III) iodide -->
Molecular equation
Total ionic
Net ionic
4] Potassium dichromate + ammonium cyanide -->
Molecular equation
Total ionic
Net ionic
5] Phosphoric acid + silver hydroxide -->
Molecular equation
Total ionic
Net ionic

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6] Sodium fluoride + hydrochloric acid -->

Molecular equation

Total ionic

Net ionic

7] Barium hydroxide + cupric acetate -->

Molecular equation

Total ionic

Net ionic

8] Nickel (II) iodide + potassium chromate -->

Molecular equation

Total ionic

Net ionic

9] Zinc + hydrobromic acid -->

Molecular equation

Total ionic

Net ionic

10] Cesium carbonate + chromium (II) acetate -->

Molecular equation

Total ionic

Net ionic
11] Silver hydroxide + nitric acid -->
Molecular equation
Total ionic
Net ionic
12] Iron (III) sulfate + sodium phosphate -->
Molecular equation
Total ionic
Net ionic
13] Chromous nitrate + sodium chloride -->
Molecular equation
Total ionic
Net ionic
14] Lithium sulfide + cobalt (II) acetate -->
Molecular equation
Total ionic
Net ionic
15] Phosphoric acid + lead (II) acetate -->
Molecular equation
Total ionic
Net ionic

