NAMES OF SIMPLE IONS

Cations:

	-OUS		-IC
Cr^{2+}	chromium (II) chromous	Cr^{3+}	chromium (III) chromic
Mn^{2+}	manganese (II) manganous	Mn^{3+}	manganese (III) manganic
Fe^{2+}	iron (II) ferrous	Fe^{3+}	iron (III) ferric
Co^{2+}	cobalt (II) cobaltous	Co^{3+}	cobalt (III) cobaltic
Ni^{2+}	nickel (II) nickelous	Ni^{3+}	nickel (III) nickelic
Cu^{+}	copper (I) cuprous	Cu^{2+}	copper (II) cupric
Au^{+}	gold (I) aurous	Au^{3+}	gold (III) auric
Hg_2^{2+}	mercury (I) mercurous	Hg^{2+}	mercury (II) mercuric
Sn^{2+}	tin (II) stannous	Sn^{4+}	tin (IV) stannic
Pb^{2+}	lead (II) plumbous	Pb^{4+}	lead (IV) plumbic
As^{3+}	arsenic (III)	As^{5+}	arsenic (V)
Sb^{3+}	antimony (III)	Sb^{5+}	antimony (V)
Bi^{3+}	bismuth (III)	Bi^{5+}	bismuth (V)

Anions:

	-IDE
N^{3-}	nitride
O^{2-}	oxide
S^{2-}	sulfide
H^{-}	hydride
F^{-}	fluoride
Cl^{-}	chloride
Br^{-}	bromide
I^{-}	iodide

For all other cations, the name of the ion is the same as the name of the element from which it derives.

Simple Cations and Anions

H^+																(H^-)	
Li^+	Be^{2+}													N^{3-}	O^{2-}	F^-	
Na^+	Mg^{2+}											Al^{3+}		P^{3-}	S^{2-}	Cl^-	
K^+	Ca^{2+}				Cr^{2+} Cr^{3+}	Mn^{2+} Mn^{3+}	Fe^{2+} Fe^{3+}	Co^{2+} Co^{3+}	Ni^{2+} Ni^{3+}	Cu^+ Cu^{2+}	Zn^{2+}			As^{3+} As^{5+}	Se^{2-}	Br^-	
Rb^+	Sr^{2+}									Ag^+	Cd^{2+}		Sn^{2+} Sn^{4+}	Sb^{3+} Sb^{5+}	Te^{2-}	I^-	
Cs^+	Ba^{2+}									Au^+ Au^{3+}	Hg^{2+} Hg_2^{2+}		Pb^{2+} Pb^{4+}	Bi^{3+} Bi^{5+}			

POLYATOMIC IONS

ATIONS

NH_4^+	ammonium

ANIONS

-1		-2		-3	
HSO_3^-	bisulfite	SO_3^{2-}	sulfite		
HSO_4^-	bisulfate	SO_4^{2-}	sulfate		
		$S_2O_3^{2-}$	thiosulfate		
HCO_3^-	bicarbonate	CO_3^{2-}	carbonate		
HS^-	bisulfide			PO_3^{3-}	phosphite
$H_2PO_4^-$	dihydrogen phosphate	HPO_4^{2-}	monohydrogen phosphate	PO_4^{3-}	phosphate
CN^-	cyanide			AsO_4^{3-}	arsenate
SCN^-	thiocyanate	CrO_4^{2-}	chromate	BO_3^{3-}	borate
OCN^-	cyanate	$Cr_2O_7^{2-}$	dichromate		
NO_2^-	nitrite				
NO_3^-	nitrate	$C_2O_4^{2-}$	oxalate		
ClO^-	hypochlorite				
ClO_2^-	chlorite	O_2^{2-}	peroxide		
ClO_3^-	chlorate				
ClO_4^-	perchlorate				
BrO^-	hypobromite				
BrO_2^-	bromite				
BrO_3^-	bromate				
BrO_4^-	perbromate				
IO^-	hypoiodite				
IO_2^-	iodite				
IO_3^-	iodate				
IO_4^-	periodate				
MnO_4^-	permanganate				
OH^-	hydroxide				
$C_2H_3O_2^-$	acetate				

PREFIXES AND SUFFIXES (what they mean)

-ate	"most common variety"	-ide	only one kind of atom in the anion
-ite	one less oxygen atom than "ate" variety (same charge)	thio-	one oxygen atom replaced by S
per-	one more oxygen atom than in "ate" variety (same charge)	bi-	one H^+ added to divalent anion
hypo-	one less oxygen atom than in "ite" variety (same charge)	di-	two