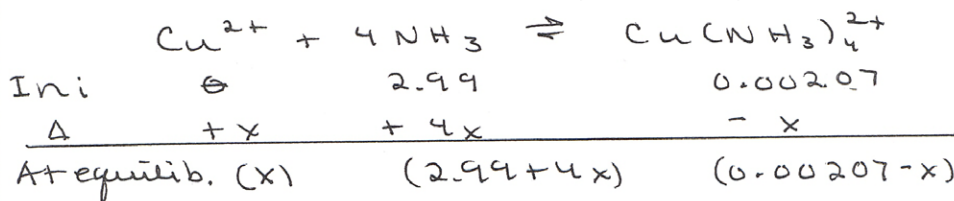
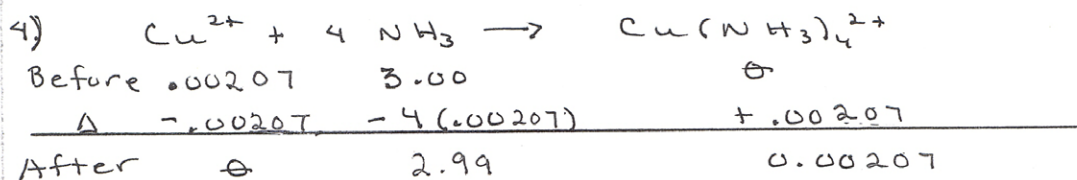


$$\begin{aligned}
 K_b &= 5. \times 10^9 = \frac{[\text{Co(OH)}_4^{2-}]}{[\text{Co}^{2+}][\text{OH}^-]^4} \\
 &= \frac{(0.0236-x)^{\rightarrow \text{neg}}}{(x)(0.706-4x)^{\rightarrow \text{neg}}^4}
 \end{aligned}$$

$$[\text{Co}^{2+}] = x = 1.9 \times 10^{-11} \text{ M}$$



$$\begin{aligned}
 K_b &= 5.6 \times 10^{11} = \frac{[\text{Cu(NH}_3)_4^{2+}]}{[\text{Cu}^{2+}][\text{NH}_3]^4} = \frac{(0.00207-x)^{\rightarrow \text{neg}}}{(x)(2.99+4x)^{\rightarrow \text{neg}}^4} \\
 [\text{Cu}^{2+}] &= x = 4.6 \times 10^{-17} \text{ M}
 \end{aligned}$$