

Net-Ionic Equations
CHEMISTRY 110

Name _____
Last _____ First _____

- a) Write a balanced equation for the reactants given.
- b) Include the physical states for all reagents: Assume that all reactions are in water.
- c) Write a total ionic equation
- d) Write the net-ionic equation

hint: Use solubility rules, activity tables, and tables for strong bases and acids to write the equations!



Molecular equation _____

Total ionic _____

Net ionic _____



Molecular equation _____

Total ionic _____

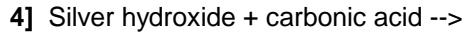
Net ionic _____



Molecular equation _____

Total ionic _____

Net ionic _____



Molecular equation _____

Total ionic _____

Net ionic _____



Molecular equation _____

Total ionic _____

Net ionic _____



Molecular equation _____

Total ionic _____

Net ionic _____

7] Cesium cyanide + hydrochloric acid -->

Molecular equation _____

Total ionic _____

Net ionic _____

8] Barium hydroxide + cupric acetate -->

Molecular equation _____

Total ionic _____

Net ionic _____

9] Chromium (III) chloride + sodium nitrate -->

Molecular equation _____

Total ionic _____

Net ionic _____

10] Iron + nickel(III) iodide -->

Molecular equation _____

Total ionic _____

Net ionic _____

11] Sodium hypoiodite + ammonium dichromate -->

Molecular equation _____

Total ionic _____

Net ionic _____

12] Acetic acid + Chromic hydroxide -->

Molecular equation _____

Total ionic _____

Net ionic _____

13] Ferrous nitrate + sodium sulfide -->

Molecular equation _____

Total ionic _____

Net ionic _____

14] Phosphoric acid + cobalt (II) bromide -->

Molecular equation _____

Total ionic _____

Net ionic _____

15] Sodium sulfide + lead (II) nitrate-->

Molecular equation _____

Total ionic _____

Net ionic _____