

EXERCISE 6

Chem 100

(Due date _____)

10 points

Name _____

(last)

(first)

LectureSection # _____ Instructor _____

Below is a list of salts and covalent compounds. In the blank in column A, write "S" if the compound is soluble in water and "I" if the compound is not soluble in water. (To determine whether a given salt is soluble in water, consult your solubility rules for salts. To determine whether a given covalent compound is soluble in water you need to know if it is polar or nonpolar. Therefore, covalent compounds below are labeled to tell you that they are polar or nonpolar.)

In column B you are to indicate whether the given compound exists in aqueous (water) solution as ions, "I", or molecules, "M". If the compound is not soluble, the blank in column B should be left empty.

	A	B
1. Ag_3PO_4	_____	_____
2. KHSO_3	_____	_____
3. CaI_2	_____	_____
4. CCl_4 (nonpolar)	_____	_____
5. $\text{Mg}(\text{NO}_2)_2$	_____	_____
6. CuCl_2	_____	_____
7. NH_4ClO_2	_____	_____
8. $\text{C}_6\text{H}_{12}\text{O}_6$ (polar)	_____	_____
9. ZnS	_____	_____
10. PbSO_4	_____	_____
11. FeCl_3	_____	_____
12. AgBr	_____	_____
13. $\text{Co}(\text{C}_2\text{H}_3\text{O}_2)_3$	_____	_____
14. NiSO_4	_____	_____
15. C_6H_6 (nonpolar)	_____	_____
16. CuF_2	_____	_____
17. $\text{Ba}(\text{NO}_3)_2$	_____	_____
18. Na_2CO_3	_____	_____
19. $\text{C}_2\text{H}_6\text{O}$ (polar)	_____	_____

20. $\text{Li}_2\text{C}_2\text{O}_4$
