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Chem 100 (Due in lab ______) 10 points

Name		
(last)		(first)
Lab Section #	Lab Instructor	

METRIC-METRIC CONVERSIONS

- How many centigrams are in 7,586 micrograms?
- 0.7586

How many liters are in 1.5 milliliters?

2. 0.0015

How many micrometers are in 0.049 decimeters? 3.

3. 4900

How many millimeters are in 43.29 centimeters?

4. 432.9

How many kilograms are in 2,080 grams?

5. 2.080

How many deciliters are in 0.289 kiloliters?

6.__ 2890

How many centigrams are in 12 decigrams?

7. 120

How many meters are in 596 centimeters?

- 8.___ 5.96
- ENGLISH-METRIC CONVERSIONS. Use English-Metric conversion factors provided in your lecture outline to solve the following. (Some English abbreviations you might need: ounce = oz., pound = lb. quart = qt. mile = mi. foot = ft. yard = yd.) **SHOW WORK**
 - How many grams are in 0.20 ounces?

How many cubic inches are in 328 milliliters?

328 mL
$$\left(\frac{1 \text{ in}^3}{16.4 \text{ mL}}\right) = 20 \text{ in}^3$$

3. How many kilograms are in 15.4 pounds?

4. How many quarts are in 4.5 liters? 4.5 L (1069t) = 4.779t

- 5. 800 m²
- 960 yd² $\left(\frac{1 \text{ m}^2}{1.20 \text{ yd}^2}\right) = 800 \text{ m}^2$

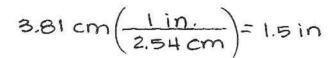
6. How many kilometers are in 28 miles?

7. How many feet are in 15 meters?

8. How many inches are in 3.81 centimeters?

20	1	S	
8	1.5	10	

49.2 ft



C. Complete the following table.

°F	- ℃	K	
260,6	127	400	annear
98.6	37	310	00:050
-40	-40	233	
59	15	288	
5	-15	258	
80	26,7	299,7	essión.

Solve the following density problems.

SHOW WORK

40.0 mL of a substance has a mass of 30.0 grams.
What is the density of the substance?

- $\frac{30.0q}{40.0 \text{ mL}} = 0.75 \text{ g/mL}$
- Substance A has a density of 4.0 g/mL. What volume (in mL) would 12.0 grams of A occupy?

- $12.0g\left(\frac{1mL}{4.0g}\right) = 3.0 mL$
- 3. Lead has a density of 11.3 g/mL. What is the mass (in grams) of a lead block that is 8 cm long by 5 cm wide by 2 cm high?