

## Common Polyatomic Ions

(Alphabetical order by ion name)

**NOTE:** *-ite* ending means one less oxygen than the *-ate* form.

Ion Name	Ion Symbol	Ion Name	Ion Symbol	Ion Name	Ion Symbol	Ion Name	Ion Symbol
Acetate	CH <sub>3</sub> CO <sub>2</sub> <sup>-</sup> or CH <sub>3</sub> COO <sup>-</sup>	Dichromate	Cr <sub>2</sub> O <sub>7</sub> <sup>2-</sup>	Hypobromite	BrO <sup>-</sup>	Perrhenate	ReO <sub>4</sub> <sup>-</sup>
Ammonium	NH <sub>4</sub> <sup>+</sup>	Dihydrogen Phosphate	H <sub>2</sub> PO <sub>4</sub> <sup>-</sup>	Hypochlorite	ClO <sup>-</sup>	Phosphate	PO <sub>4</sub> <sup>3-</sup>
Arsenate	AsO <sub>4</sub> <sup>3-</sup>	Dihydrogen Phosphite	H <sub>2</sub> PO <sub>3</sub> <sup>-</sup>	Hypoiodite	IO <sup>-</sup>	Phosphite	PO <sub>3</sub> <sup>3-</sup>
Arsenite	AsO <sub>3</sub> <sup>3-</sup>	Dithionate	S <sub>2</sub> O <sub>6</sub> <sup>2-</sup>	Iodate	IO <sub>3</sub> <sup>-</sup>	Plumbate	PbO <sub>3</sub> <sup>2-</sup>
Azide	N <sub>3</sub> <sup>-</sup>	Dithionite	S <sub>2</sub> O <sub>4</sub> <sup>2-</sup>	Iodite	IO <sub>2</sub> <sup>-</sup>	Plumbite	PbO <sub>2</sub> <sup>2-</sup>
Borate	BO <sub>3</sub> <sup>2-</sup>	Ferricyanide	Fe(CN) <sub>6</sub> <sup>3-</sup>	Isocyanate	NCO <sup>-</sup>	Rhenate	ReO <sub>4</sub> <sup>2-</sup>
Bromate	BrO <sub>3</sub> <sup>-</sup>	Ferrocyanide	Fe(CN) <sub>6</sub> <sup>4-</sup>	Manganate	MnO <sub>4</sub> <sup>2-</sup>	Selenate	SeO <sub>4</sub> <sup>2-</sup>
Bromite	BrO <sub>2</sub> <sup>-</sup>	Fulminate	CNO <sup>-</sup>	Nitrate	NO <sub>3</sub> <sup>-</sup>	Selenite	SeO <sub>3</sub> <sup>2-</sup>
Carbonate	CO <sub>3</sub> <sup>2-</sup>	Hydrazide	N <sub>2</sub> H <sub>3</sub> <sup>-</sup>	Nitrite	NO <sub>2</sub> <sup>-</sup>	Silicate	SiO <sub>3</sub> <sup>2-</sup>
Chlorate	ClO <sub>3</sub> <sup>-</sup>	Hydrogen arsenate	HAsO <sub>4</sub> <sup>2-</sup>	Oxalate	C <sub>2</sub> O <sub>4</sub> <sup>2-</sup>	Stannate	SnO <sub>3</sub> <sup>2-</sup>
Chlorite	ClO <sub>2</sub> <sup>-</sup>	Hydrogen carbonate	HCO <sub>3</sub> <sup>-</sup>	Ozonide	O <sub>3</sub> <sup>-</sup>	Stannite	SnO <sub>2</sub> <sup>2-</sup>
Chromate	CrO <sub>4</sub> <sup>2-</sup>	Hydrogen Phosphate	HPO <sub>4</sub> <sup>2-</sup>	Perbromate	BrO <sub>4</sub> <sup>-</sup>	Sulfate	SO <sub>4</sub> <sup>2-</sup>
Chromite	CrO <sub>2</sub> <sup>-</sup>	Hydrogen Phosphite	HPO <sub>3</sub> <sup>2-</sup>	Perchlorate	ClO <sub>4</sub> <sup>-</sup>	Sulfite	SO <sub>3</sub> <sup>2-</sup>
Citrate	C <sub>6</sub> H <sub>5</sub> O <sub>7</sub> <sup>3-</sup>	Hydrogen Sulfate	HSO <sub>4</sub> <sup>-</sup>	Periodate	IO <sub>4</sub> <sup>-</sup>	Superoxide	O <sub>2</sub> <sup>-</sup>
Cyanate	OCN <sup>-</sup>	Hydrogen Sulfite	HSO <sub>3</sub> <sup>-</sup>	Permanganate	MnO <sub>4</sub> <sup>-</sup>	Thiocyanate	SCN <sup>-</sup>
Cyanide	CN <sup>-</sup>	Hydroxide	OH <sup>-</sup>	Peroxide	O <sub>2</sub> <sup>2-</sup>	Thiosulfate	S <sub>2</sub> O <sub>3</sub> <sup>2-</sup>
						Tungstate	WO <sub>4</sub> <sup>2-</sup>

### PREFIXES:

*per-* = one more oxygen than *-ate*

*hypo-* = one less oxygen than *-ite*

## Common Polyatomic Ions

(Alphabetical order by formula)

Ion Symbol	Ion Name	Ion Symbol	Ion Name	Ion Symbol	Ion Name	Ion Symbol	Ion Name
$\text{AsO}_3^{3-}$	Arsenite	$\text{C}_2\text{O}_4^{2-}$	Oxalate	$\text{IO}_3^-$	Iodate	$\text{PO}_3^{3-}$	Phosphite
$\text{AsO}_4^{3-}$	Arsenate	$\text{CrO}_2^-$	Chromite	$\text{IO}_4^-$	Periodate	$\text{PO}_4^{3-}$	Phosphate
$\text{BO}_3^{2-}$	Borate	$\text{CrO}_4^{2-}$	Chromate	$\text{MnO}_4^{2-}$	Manganate	$\text{PbO}_2^{2-}$	Plumbite
$\text{BrO}^-$	Hypobromite	$\text{Cr}_2\text{O}_7^{2-}$	Dichromate	$\text{MnO}_4^-$	Permanganate	$\text{PbO}_3^{2-}$	Plumbate
$\text{BrO}_2^-$	Bromite	$\text{Fe}(\text{CN})_6^{3-}$	Ferricyanide	$\text{N}_3^-$	Azide	$\text{ReO}_4^{2-}$	Rhenate
$\text{BrO}_3^-$	Bromate	$\text{Fe}(\text{CN})_6^{4-}$	Ferrocyanide	$\text{NCO}^-$	Isocyanate	$\text{SCN}^-$	Thiocyanate
$\text{BrO}_4^-$	Perbromate	$\text{HAsO}_4^{2-}$	Hydrogen arsenate	$\text{NH}_2^-$	Amide	$\text{SO}_3^{2-}$	Sulfite
$\text{CH}_3\text{CO}_2^-$ or $\text{CH}_3\text{COO}^-$	Acetate	$\text{HCO}_3^-$	Hydrogen carbonate	$\text{NH}_4^+$	Ammonium	$\text{SO}_4^{2-}$	Sulfate
$\text{C}_6\text{H}_5\text{O}_7^{3-}$	Citrate	$\text{HPO}_3^{2-}$	Hydrogen Phosphite	$\text{N}_2\text{H}_3^-$	Hydrazide	$\text{S}_2\text{O}_3^{2-}$	Thiosulfate
$\text{ClO}^-$	Hypochlorite	$\text{HPO}_4^{2-}$	Hydrogen Phosphate	$\text{NO}_2^-$	Nitrite	$\text{S}_2\text{O}_4^{2-}$	Dithionite
$\text{ClO}_2^-$	Chlorite	$\text{H}_2\text{PO}_3^-$	Dihydrogen Phosphite	$\text{NO}_3^-$	Nitrate	$\text{S}_2\text{O}_6^{2-}$	Dithionate
$\text{ClO}_3^-$	Chlorate	$\text{H}_2\text{PO}_4^{2-}$	Dihydrogen Phosphate	$\text{O}_2^-$	Superoxide	$\text{SeO}_3^{2-}$	Selenite
$\text{ClO}_4^-$	Perchlorate	$\text{HSO}_3^-$	Hydrogen Sulfite	$\text{O}_2^{2-}$	Peroxide	$\text{SeO}_4^{2-}$	Selenate
$\text{CN}^-$	Cyanide	$\text{HSO}_4^-$	Hydrogen Sulfate	$\text{O}_3^-$	Ozonide	$\text{SiO}_3^{2-}$	Silicate
$\text{CNO}^-$	Fulminate	$\text{IO}^-$	Hypoiodite	$\text{OCN}^-$	Cyanate	$\text{SnO}_2^{2-}$	Stannite
$\text{CO}_3^{2-}$	Carbonate	$\text{IO}_2^-$	Iodite	$\text{OH}^-$	Hydroxide	$\text{SnO}_3^{2-}$	Stannate
						$\text{WO}_4^{2-}$	Tungstate

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